

# Theoretical Study of Acetylation of Ethylamine catalyzed by Zn<sup>2+</sup> ions

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A theoretical study of acetylation of ethylamine catalyzed by Zn<sup>2+</sup> ions from the analysis of intermediate of the reaction was carried out.

The study of acetylation of amines is of great interest by the utility of its products of reaction and is one of the most frequently used transformations in organic synthesis as it provides an efficient and inexpensive means for protecting amino groups in a multistep synthetic process

Acetylation of amine is a nucleophilic substitution reaction. This reaction can be catalyzed by Lewis acid, metallic ion. In reaction mechanism, the metallic ion formed a complex with the oxygen of the acetic anhydride carbonyl, facilitating the polarization of the same and the successive addition of amine at the position to form a tetrahedral intermediate, determining step of the rate of the reaction.

In the present theoretical work were investigated the structure and energy of the tetrahedral intermediate of the reaction catalyzed by Zn<sup>2+</sup> ions. Geometries of all species involved in the acetylation were made and identified. All of the geometry optimizations were performed by the method at the DFT/B3LYP level of theory and the method MP2. Were adopted the 6-31+G\* basis sets. Energies were calculated using the Mechanics-UFF method. Following the same procedure it was identified the geometric parameters and energy of reaction intermediate.

The calculations show 57.82 kcal/mol of energy for the tetrahedral intermediate and the energy of activation for the reaction was 9.59 kcal/mol.